

Colligative Properties Of Solutions Section Review

Answers

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SECTION 16.3 COLLIGATIVE PROPERTIES OF SOLUTIONS pages 487–490 This section explains why a solution has a lower vapor pressure an elevated boiling point and a depressed freezing point compared with the pure solvent of that solution

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CHAPTER 16 Colligative Properties of Solutions

Practice Problem 6660020 book December 22 2010 8:37 Chapter 16 Colligative Properties of Solutions 45–16.4 The mole fraction of NH_4SO_4 aq is given by $x_{\text{NH}_4\text{SO}_4} = \frac{n_{\text{NH}_4\text{SO}_4}}{n_{\text{NH}_4\text{SO}_4} + n_{\text{H}_2\text{O}}}$ Because NH_4SO_4 aq is a strong electrolyte it dissociates completely into NH_4^+ aq and SO_4^{2-} aq ions Assume a one kilogram solution

Solutions and Colligative Properties Saptarshi Classes

Colligative properties The properties of solutions that depend only on the number of solute particles in solution and not on the nature of the solute particle are called as Colligative properties SAPTARSHI

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of the colligative properties Osmotic pressure provides the most accurate determination because of the magnitude of $\hat{1}$ 0.0200 M solution of glucose exerts osmotic pressure of 374 mm Hg 0.5 atm but freezing point depression of only 0.02°C

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Colligative Properties Page 1 Lecture 4 Colligative

Colligative Properties Page 1 Lecture 4 Colligative Properties \hat{c} By definition a colligative property is a solution property a property of mixtures for which it is the amount of solute dissolved in the solvent matters but the kind of solute does not matter

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Colligative Properties Problems Answers Remember that Colligative properties depend only on the number or concentration of particles in a solution The properties are for ideal solutions independent of the kind or size of the particles whether ionic large or small etc

Colligative Properties of Solutions \hat{c} Introductory

Two colligative properties are related to solution concentration as expressed in molality As a review recall the

definition of molality Because the vapour pressure of a solution with a nonvolatile solute is depressed compared to that of the pure solvent it requires a higher temperature for the solution's vapour pressure to reach 1 00 atm 760 torr

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Chapter 16

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