

Titration Practice Worksheet Answers

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Titration Practice Worksheet Find the requested quantities in the following problems 1 2 3 Explain your answer in a few sentences j r 5 Explain the difference between an endpoint and equivalence point in a titration 0 f c 6 How many moles of LiOH are needed to exactly neutralize 2 0 moles of H₂SO₄

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Explain your answer in a few sentences 5 Explain the difference between an endpoint and equivalence point in a titration For questions 1 and 2 the units for your final answer should be mol/L or molar because you're trying to find the molarity of the acid or base solution To solve these problems use $M_1V_1 = M_2V_2$

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Titration Practice Worksheet Find the requested quantities in the following problems 1 If it takes 54 mL of 0 1 M NaOH to neutralize 125 mL of an HCl solution answer in a few sentences 5 Explain the difference between an endpoint and equivalence point in a titration

Titration Practice Worksheet

Solutions to the Titration Practice Worksheet For questions 1 and 2 the units for your final answer should be mol/L or molar because you're trying to find the molarity of the acid or base solution To solve these problems use $M_1V_1 = M_2V_2$

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The titration curve in Figure 1 shows a strong acid being titrated by a strong base There is the initial slow rise in pH until the reaction nears the point where just enough base is added to neutralize all the initial acid This point is the equivalence point Use the graph and reading to answer the questions below on a separate sheet of paper

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Answers to Question 6b

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Titration worksheet W 336 Everett Community College Tutoring Center Student Support Services Program 1
It takes 83 mL of a 0.45 M NaOH solution to neutralize 235 mL of an HCl solution. What is the concentration of the HCl solution? 2 You are titrating an acid into a base to determine the concentration of the base. The

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Titration Problems 1 A 0.15 M solution of NaOH is used to titrate 200 mL of 0.15 M HCN. What is the pH at the equivalence point? $K_a = 4.9 \times 10^{-10}$ 2 A 0.25 M solution of HCl is used to titrate 0.25 M NH_3 . What is the pH at the

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